



# Cu<sub>2</sub>O/TiO<sub>2</sub> heterostructures for CO<sub>2</sub> reduction through a direct Z-scheme: Protecting Cu<sub>2</sub>O from photocorrosion

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## ABSTRACT

The development of artificial photosynthesis aims to solve the increasing energy demand and associated environmental problems. A model photosynthetic system employing a composite of semiconductors with a Z-scheme can potentially mimic the combined power of photosystems 1 and 2 to transfer electrons. In this work, octahedral cuprous oxide covered with titanium dioxide nanoparticles (Cu<sub>2</sub>O/TiO<sub>2</sub>) are synthesized by a solvothermal strategy that provides high morphological and crystallographic control. The formation of a *p-n* heterojunction and characterization of the Type II band alignment of the composite are performed by diffuse reflectance UV-visible (DRUV) spectroscopy, ultraviolet photoelectron spectroscopy (UPS), and X-ray photoelectron spectroscopy (XPS). Upon UV-visible irradiation ( $\lambda \geq 305$  nm) of the composite in the presence of water vapor as the hole scavenger, the photoreduction of CO<sub>2</sub>(g) proceeds selectively to generate CO(g). The production rate of CO by the composite,  $R_{CO} = 2.11 \mu\text{mol g}_{\text{cat}}^{-1} \text{h}^{-1}$ , is 4-times larger than for pure Cu<sub>2</sub>O under identical conditions. Contrasting XPS analyses of Cu<sub>2</sub>O and Cu<sub>2</sub>O/TiO<sub>2</sub>, during photocatalysts operation and the detection of photogenerated hydroxyl radicals (HO•) in the heterostructure at variance with the results obtained for pure Cu<sub>2</sub>O are taken as evidences that TiO<sub>2</sub> protects Cu<sub>2</sub>O from undergoing photocorrosion. These results provide direct evidence of an efficient Z-scheme as the main mechanism for harvesting energy during CO<sub>2</sub> reduction in the synthesized materials.

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## 1. Introduction

The design of efficient photocatalysts aimed at the conversion of CO<sub>2</sub> to useful chemicals is a scientific challenge which is driven by the shortage of energy resources and the increased concentration of this greenhouse gas in the atmosphere [1–4]. The high stability of the linear CO<sub>2</sub> molecule makes its reduction difficult and usually demands scarce and expensive species, such as ruthenium or rhenium complexes [4–7]. The use of heterogeneous photocatalytic systems involving TiO<sub>2</sub>, ZnO, or ZnS provide a possible, although less efficient alternative for carbon dioxide reduction under UV excitation [8–10]. It should be noticed that none of these semiconductors provides sufficient potential to mediate the one electron reduction of CO<sub>2</sub> (Equation (1)), since their conduction bands lie

below the homogeneous reduction potential of carbon dioxide,  $E^{\circ}(\text{CO}_2/\text{CO}_2^{\bullet-}) = -1.9 \text{ V}$  [11].

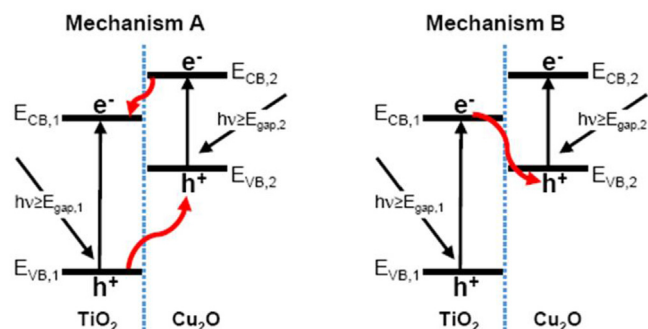


Although adsorption on semiconductor surfaces would certainly bend the linear CO<sub>2</sub> molecule and lower the energy demand for reaction 1, the considerable gap between the CO<sub>2</sub> LUMO and the conduction band (CB) probably prevents direct CO<sub>2</sub> reduction by CB electrons [12]. Thus, the photocatalytic reduction of CO<sub>2</sub> frequently involves the use of sacrificial donors (*i.e.*, alcohols) that generate reactive intermediates (hydroxyalkyl radicals) which can accomplish reaction (1). It is also apparent that the low photon intermittency usually employed under practical photocatalytic experiments is against the occurrence of multielectronic processes [13], which could by-pass the high energy intermediate CO<sub>2</sub>•<sup>-</sup> [14].

Cuprous oxide (Cu<sub>2</sub>O) is an interesting *p*-semiconductor whose conduction band is much more energetic than the commonly used wide bandgap semiconductors. The bandgap of cuprous oxide is  $E_{\text{gap},2} \approx 2.2 \text{ eV}$ , while its conduction band is poised at

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**Scheme 1.** Schematic representation of the processes that can take place in a system consisting of two semiconductors in electrical contact under irradiation with photons of suitable energy. (A) Double-charge transfer, and (B) direct Z-scheme mechanism.

$E_{CB,2} = -1.4 \text{ eV}$  vs NHE at  $\text{pH} = 7$  [15]. Unfortunately, as commonly found for semiconductors active in the visible region its photostability is scarce [16]. However, it can be envisaged that surface modification of  $\text{Cu}_2\text{O}$  with a  $n$ -type wide bandgap semiconductor such as  $\text{TiO}_2$  could generate a Type II heterostructure and avoid  $\text{Cu}_2\text{O}$  photocorrosion as discussed below.

Scheme 1 shows the expected energy diagram for the Type II  $\text{Cu}_2\text{O}/\text{TiO}_2$  heterostructure [17].

Simultaneous excitation of the individual components of the composite leads to an electron-hole pair in each photocatalytic center. Different mechanisms can be envisaged to rationalize the charge transfer across the interface in  $p$ - $n$  heterojunctions. Scheme 1A represents a double-charge transfer mechanism, in which the  $\text{Cu}_2\text{O}$  electrons move to the  $\text{TiO}_2$  center and  $\text{TiO}_2$  holes migrate to  $\text{Cu}_2\text{O}$  [18,19]. Alternatively, Scheme 1B involves a direct Z-scheme mechanism [20–23] where the  $\text{TiO}_2$  electrons are used to scavenge  $\text{Cu}_2\text{O}$  holes. The double-charge transfer mechanism favors charge separation, at the expense of a decrease in the potential energy of electrons and holes. Moreover, it generates an excess of holes in  $\text{Cu}_2\text{O}$  that could lead to its photocorrosion in absence of a suitable electron donor. Instead, under the Z-scheme mechanism the electron transfer across the interface could provide a way for preserving  $\text{Cu}_2\text{O}$  photostability while maintaining a high reduction potential. Both, mechanisms A and B (Scheme 1), have been previously invoked to explain the higher efficiency of the heterostructures in comparison to the individual counterparts; however, the factors favoring one of the two mechanisms remain elusive.

In this work, we synthesize and fully characterize a type II heterostructure based on  $p$ -type octahedral  $\text{Cu}_2\text{O}$  and  $n$ -type  $\text{TiO}_2$  nanoparticles. The synthesized material shows an enhanced efficiency for  $\text{CO}_2$  photoreduction in comparison to the individual materials. Based on comparative X-ray photoelectron spectroscopy (XPS) studies of  $\text{Cu}_2\text{O}$  and  $\text{Cu}_2\text{O}/\text{TiO}_2$  heterostructures under UV-visible irradiation, we provide direct experimental evidence in favor of the Z-scheme mechanism.

## 2. Experimental

### 2.1. Chemicals and materials

Copper(II) chloride dihydrate ( $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ , 99.4% assay, J. T. Baker), polyvinylpyrrolidone (PVP, average  $\text{MW} \approx 29000$  from Aldrich), sodium hydroxide (NaOH, 99.3% assay, VWR), L-ascorbic acid (99.7% assay, Sigma-Aldrich), ethanol (absolute for analysis, 99.9% assay, EMD Chemicals) and titanium(IV) butoxide (97.0% assay, Sigma-Aldrich) were used as received. Carbon dioxide ( $\text{CO}_2$ , UHP) and carbon monoxide (CO, CP) were purchased from Scott

Gross. All procedures employed ultrapure water ( $18.2 \text{ M}\Omega \text{ cm}$ , ELGA PURELAB flex, Veolia).

### 2.2. Synthesis of $\text{Cu}_2\text{O}$ octahedra

$\text{Cu}_2\text{O}$  synthesis closely follows a previous published procedure [24]. In a typical experiment, 6.66 g of PVP were dissolved in an aqueous solution of  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$  ( $0.01 \text{ mol L}^{-1}$ , 100 mL) at  $55^\circ\text{C}$ . Then, 10.0 mL NaOH aqueous solution ( $2.0 \text{ mol L}^{-1}$ ) was added dropwise into the above transparent light green solution. During this process, the solution color changes from blue-green to dark brown. The reaction mixture was stirred for 0.5 h, keeping the temperature at  $55^\circ\text{C}$ . Afterwards, 10 mL of an aqueous ascorbic acid solution ( $0.6 \text{ mol L}^{-1}$ ) was added dropwise and the mixture was aged for 3 h under constant stirring and strict temperature control ( $55^\circ\text{C}$ ). This last stage gradually produces a red suspension. The resulting precipitate was collected by centrifugation and washed consecutively with 10 mL of deionized water and absolute ethanol (5 times). Finally, the solid was dried under vacuum at  $60^\circ\text{C}$  for 5 h for further use and characterization.

When mentioned a thermal treatment was applied to the as-synthesized solid. Briefly, the solid was resuspended in absolute ethanol, transferred to 200 mL Teflon-lined stainless steel autoclave and heated at  $180^\circ\text{C}$  for 12 h in a programmable oven. An initial ramp of  $1^\circ\text{C min}^{-1}$  was used to achieve the final temperature.

### 2.3. Preparation of $\text{TiO}_2/\text{Cu}_2\text{O}$

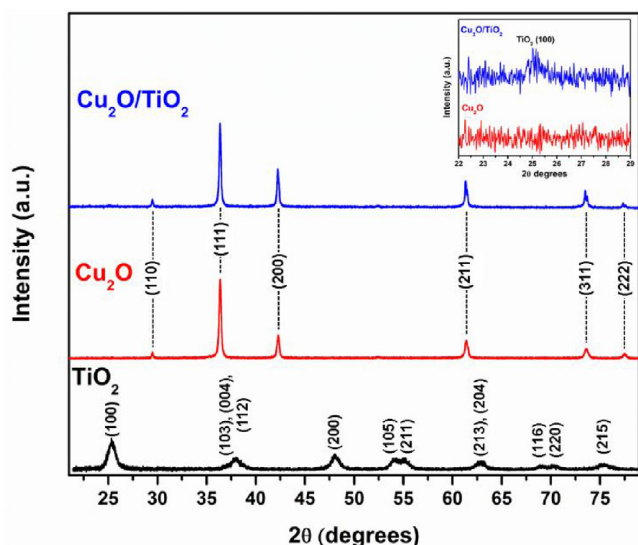
The modification of  $\text{Cu}_2\text{O}$  octahedral was performed following the procedure developed by Liu et al. [25]. Briefly, 186 mg of  $\text{Cu}_2\text{O}$  were resuspended in 65 mL of absolute ethanol with the aid of ultrasonication for 30 min, and the suspension was cooled at  $0^\circ\text{C}$ . At this temperature, 1.3 mL of titanium (IV) butoxide ethanolic solution ( $0.1 \text{ mol L}^{-1}$ ) were added dropwise into the  $\text{Cu}_2\text{O}$  suspension and stirred for 1 h at  $0^\circ\text{C}$ . Afterwards, 6.5 mL of a water:ethanol solution (volume ratio 1:4) were added dropwise to the mixture under vigorous stirring for another hour. The reaction mixture was diluted with 98 mL of absolute ethanol, transferred to 200 mL Teflon-lined stainless steel autoclave and heated at  $180^\circ\text{C}$  for 12 h in a programmable oven. An initial ramp of  $1^\circ\text{C min}^{-1}$  was used to achieve the final temperature. The amount of  $\text{Cu}_2\text{O}$  and  $\text{TiO}_2$  precursors used in this approach guarantee a nominal proportion of 0.05:0.95 of  $\text{TiO}_2:\text{Cu}_2\text{O}$  in weight. The product was collected by centrifugation and washed consecutively with 10 mL of deionized water and absolute ethanol (5 times). Finally, the solid was dried in vacuum at  $60^\circ\text{C}$  for 5 h for further use and characterization.

Pure  $\text{TiO}_2$  was obtained using the same solvothermal procedure described above, in absence of  $\text{Cu}_2\text{O}$  crystals.

### 2.4. Catalyst characterization

The crystalline properties of the as-prepared samples were analyzed via powder X-ray diffraction (XRD). The analysis was carried out on a X'Pert PRO (PANalytical) powder X-ray diffractometer, with  $\text{Cu K}\alpha$  ( $1.54 \text{ \AA}$ ) as the incident radiation and operated at an accelerating voltage of 40 kV with a current intensity of 40 mA.

The morphology of the samples was observed by scanning electron microscopy (SEM) using a Hitachi S-4300 instrument with an accelerating voltage of 25 kV. A ca.  $10 \mu\text{L}$  drop of colloidal suspension in ethanol ( $10 \text{ mg mL}^{-1}$ ) was put on a SEM grid and dried under a red light lamp. The phase distribution in the Au-metallized samples, was analyzed by Energy-Dispersive X-ray Spectroscopy (EDS) using a Jeol JSM-6460LV scanning electron microscope, operating at 15 kV coupled to a EDS probe Genesis XM4-Sys 60. X-ray photoelectron spectroscopy (XPS) measurements of the powdered samples were conducted using a PHI VersaProbe II spectrometer with an Al



**Fig. 1.** X-ray diffraction patterns of (blue)  $\text{Cu}_2\text{O}/\text{TiO}_2$ , (red)  $\text{Cu}_2\text{O}$  and (black)  $\text{TiO}_2$ . The crystal planes from JCPDS are shown between parentheses. Inset shows an enlarged view of the region between  $\theta = 22^\circ$ – $29^\circ$  of (blue)  $\text{Cu}_2\text{O}/\text{TiO}_2$  and (red)  $\text{Cu}_2\text{O}$  diffractograms. (For interpretation of the references to colour in this figure legend, the reader is referred to the web version of this article.)

K- $\alpha$  anode (1486.6 eV photon energy, 86.6 W). Ultraviolet photoelectron spectroscopy (UPS) studies were carried out by using a PHI 5600 system with He (10.2 eV) as monochromatic light source with a polarization potential (bias) of  $-5.0$  eV. All values determined from UPS analysis are referred to vacuum. The diffuse reflectance UV–visible (DRUV) spectra of the powdered samples were obtained with an Evolution 220, ISA-220 accessory, Thermo Scientific UV–vis spectrophotometer using a built-in 10 mm silicon photodiode with a 60 mm Spectralon sphere. The ISA-220 accessory was used in a configuration to register the diffuse reflectance spectrum of dry solid powders as a Kubelka–Munk function against the certified Spectralon standard [14].

### 2.5. Photocatalytic studies

The photocatalytic experiments were performed in a  $135\text{ cm}^3$  customized quartz photoreactor with flat circular windows (diameter =  $5.08\text{ cm}$ ). Based on the elemental distribution obtained by SEM–EDS analysis (see Supporting Information, S1),  $3\text{ mL}$  of suspensions of appropriate concentrations to deposit  $30\text{ mg}$  of  $\text{Cu}_2\text{O}/\text{TiO}_2$ ,  $28.4\text{ mg}$  of  $\text{Cu}_2\text{O}$ , or  $1.6\text{ mg}$   $\text{TiO}_2$  were uniformly deposited in one of the reactor windows, and left to dry overnight. The reactor was then filled with  $1\text{ atm}$   $\text{CO}_2(\text{g})$  saturated with water vapor by sparging the gas during  $20\text{ min}$  ( $0.5\text{ L min}^{-1}$ ) through a gas wash bottle. The reactor sealed with septa was kept in the dark for  $1\text{ h}$  before irradiation to ensure a homogenous internal atmosphere. UV–visible irradiation was performed with a collimated  $1\text{ kW}$  high-pressure Hg (Xe) arc lamp provided with a water filter and a cut-off filter at  $\lambda \geq 305\text{ nm}$  (Newport) previously described [14]. The irradiance of the lamp was measured in the interval  $305 \leq \lambda \leq 665\text{ nm}$  with a calibrated radiometer (Ocean Optics). Gas aliquots of  $0.5\text{ cm}^3$  were taken from the reactor at different irradiation times for analysis by gas chromatography (SRI 8610C, Multiple Gas #3 GC) equipped with two columns (a silica gel HaySep D as column 1 and a Mole-Sieve 13X as column 2), a thermal conductivity detector (TCD), and a flame ionization (FID) detector interfaced to a methanizer. The irradiation of the photocatalyst and quantification of produced  $\text{CO}(\text{g})$  were performed at room temperature ( $20^\circ\text{C}$ ). Additionally, the reaction products were identified by FTIR spectroscopy using a  $2.4\text{ m}$  path length infrared gas cell with ZnSe windows (PIKE)

thermostatted at  $100^\circ\text{C}$  mounted in an iZ10 FTIR module connected to an infrared microscope (Thermo Scientific Nicolet iN10) [26]. Furthermore, photocatalyst alteration during irradiation was monitored by XPS using a Thermo-Scientific K-Alpha X-ray Photoelectron Spectrometer with an Al K- $\alpha$  anode ( $1486.6\text{ eV}$  photon energy,  $300\text{ W}$ ).

The formation of hydroxyl radicals ( $\text{HO}^\bullet$ ) on the surface of irradiated  $\text{Cu}_2\text{O}/\text{TiO}_2$  was quantified using the coumarin fluorescence method with a Lumina Fluorescence Spectrometer (Thermo Scientific) using excitation at  $\lambda_{\text{exc}} = 332\text{ nm}$ . While coumarin is a poorly fluorescent molecule, it is an excellent probe to quantify  $[\text{HO}^\bullet]$  trapped in the produced 7-hydroxycoumarin. 7-Hydroxycoumarin has a characteristic fluorescent signal at  $\lambda_{\text{em}} = 456\text{ nm}$  [27], which is proportional to the formed  $[\text{HO}^\bullet]$  [28]. The experimental procedure followed the same steps of the photoreduction experiments, except that  $30\text{ mg}$  of coumarin (Alfa Aesar, 98.0%) were deposited on top of the  $30\text{ mg}$  of nanocomposite thin film for a 1:1 mass ratio [20]. Each data point corresponds to individual irradiation experiments at times of 0, 1, 2, and 3 h. Controls in the dark and with the individual components were also performed. Samples were extracted with  $51\text{ mL}$  of water [20], centrifuged at  $4400\text{ rpm}$  for  $5\text{ min}$ , and forced through a filter (Acrodisc  $0.2\text{ }\mu\text{m}$  pore size; Pall Corp.) to quantify by standard addition the 7-hydroxycoumarin (Acros, 98.5%) produced.

## 3. Results and discussion

### 3.1. Selection of $\text{Cu}_2\text{O}$ shape and composite architecture

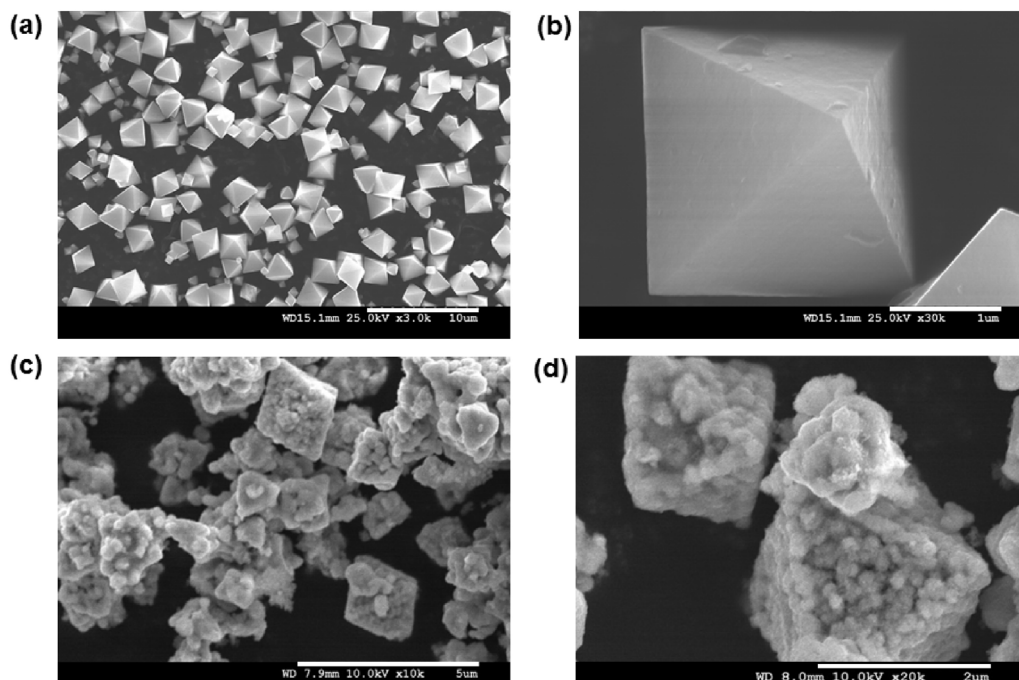
By modifying the reaction conditions, *i.e.*, the relative amount between the capping agent and  $\text{Cu}_2\text{O}$  precursors, the shape of  $\text{Cu}_2\text{O}$  crystals can be easily modified from cubic to octahedra. This issue has attracted the attention of many researchers, who concluded that octahedral  $\text{Cu}_2\text{O}$  with exposed (111) facets exhibited much higher photocatalytic activity than the cubic structure [29–31]. It is apparent that crystalline (111) facets containing active Cu atoms can increase the interaction with donors and acceptors [24,32]. Guided by these conditions, we choose  $\text{Cu}_2\text{O}$  octahedral as the starting point to design the composite [33,34]. Also, since the hybrid composite would allow free access of the acceptor and donor to both surfaces, we regulated the concentration of the semiconductor precursors taking into account the individual size and surface area in order to warrant nearly all of  $\text{TiO}_2$  nanoparticles cover the surface of  $\text{Cu}_2\text{O}$  and the amount of free  $\text{TiO}_2$  nanoparticles is depreciable as verified by SEM analysis. The actual proportion by weight in the composite was estimated by SEM–EDS analysis to be  $0.053:0.947$  of  $\text{TiO}_2:\text{Cu}_2\text{O}$ , in very good agreement with the nominal proportion (See Fig. S1).

### 3.2. XRD and SEM analysis of crystal structure and morphology

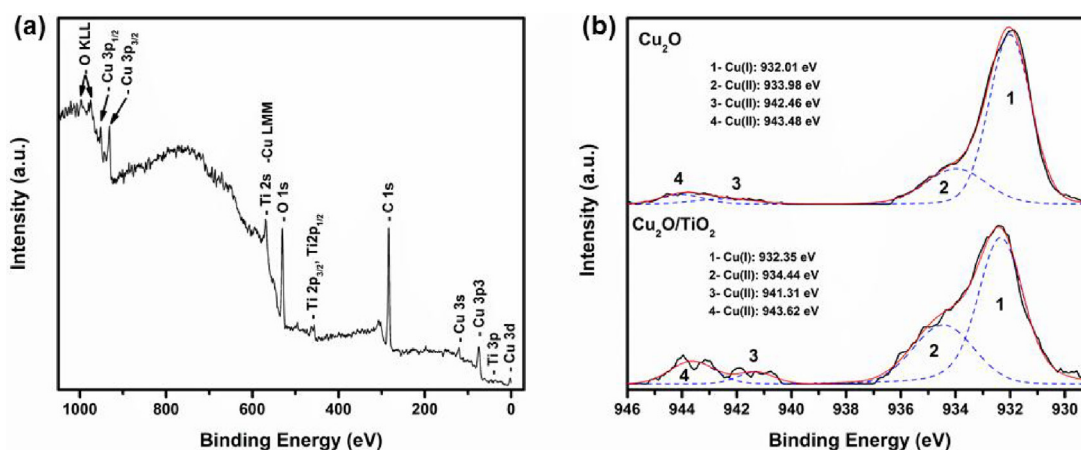
Fig. 1 shows the X-ray diffraction patterns of the synthesized  $\text{Cu}_2\text{O}$  octahedra,  $\text{TiO}_2$  and  $\text{Cu}_2\text{O}/\text{TiO}_2$ . For both  $\text{Cu}_2\text{O}$  and  $\text{Cu}_2\text{O}/\text{TiO}_2$ , all the peaks belong to the face-centered cubic  $\text{Cu}_2\text{O}$  phase (PDF Card No. 05-0667) and no diffraction peaks of CuO or metallic Cu could be detected. The strong and sharp peaks attributable to  $\text{Cu}_2\text{O}$  indicate a high degree of crystallinity. A low-intensity diffraction peak at  $\theta = 25.08$  is present in the  $\text{Cu}_2\text{O}/\text{TiO}_2$  diffraction pattern, and ascribed to the (100) plane of anatase, as shown in the inset of Fig. 1.

The XRD pattern of pure  $\text{TiO}_2$  (Fig. 1), obtained in a parallel synthesis is in agreement with the report for anatase (JCPDS 21–1272).

SEM measurements, Fig. 2(a, b), show that after solvothermal treatment of the titanium(IV) butoxide ethanolic solution in the presence of  $\text{Cu}_2\text{O}$  octahedral particles, fine  $\text{TiO}_2$  nanoparticles cover  $\text{Cu}_2\text{O}$  (Fig. 2(c, d)). It is also apparent that  $\text{Cu}_2\text{O}$  morphology is not



**Fig. 2.** SEM images for (a, b)  $\text{Cu}_2\text{O}$ , (c, d)  $\text{Cu}_2\text{O}/\text{TiO}_2$ . The left hand side panels show images obtained with low magnification, while the aspect of isolated particles obtained with high magnification are shown at the right hand side.



**Fig. 3.** (a) XPS survey spectrum for  $\text{Cu}_2\text{O}/\text{TiO}_2$  composite. (b) high resolution spectra for  $\text{Cu} 2p_{3/2}$  (peak 1) and shake-up satellite peaks (2, 3, and 4) in  $\text{Cu}_2\text{O}$  (top) and  $\text{Cu}_2\text{O}/\text{TiO}_2$  (bottom) are shown as black solid lines. For comparison, the deconvoluted components (blue dash line) and the total fit from the sum of the components (red solid line) are also shown. See text for details. (For interpretation of the references to colour in this figure legend, the reader is referred to the web version of this article.)

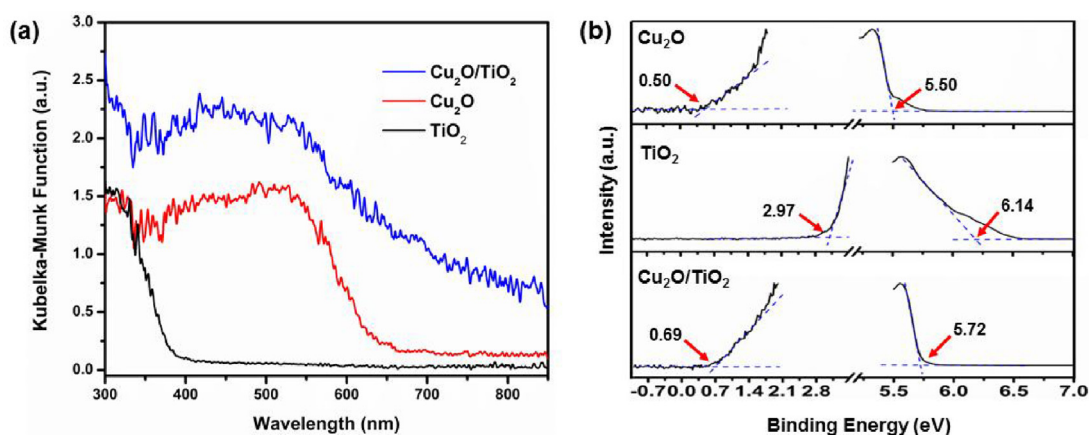
altered by the thermal treatment, thus, the difference observed in the photocatalytic activity of the composite presented below cannot be attributed to changes in the surface chemistry of  $\text{Cu}_2\text{O}$  particles, but rather to the presence of  $\text{TiO}_2$ . Distribution phase was confirmed by EDS (Fig. S1).

### 3.3. Surface characterization by XPS analysis

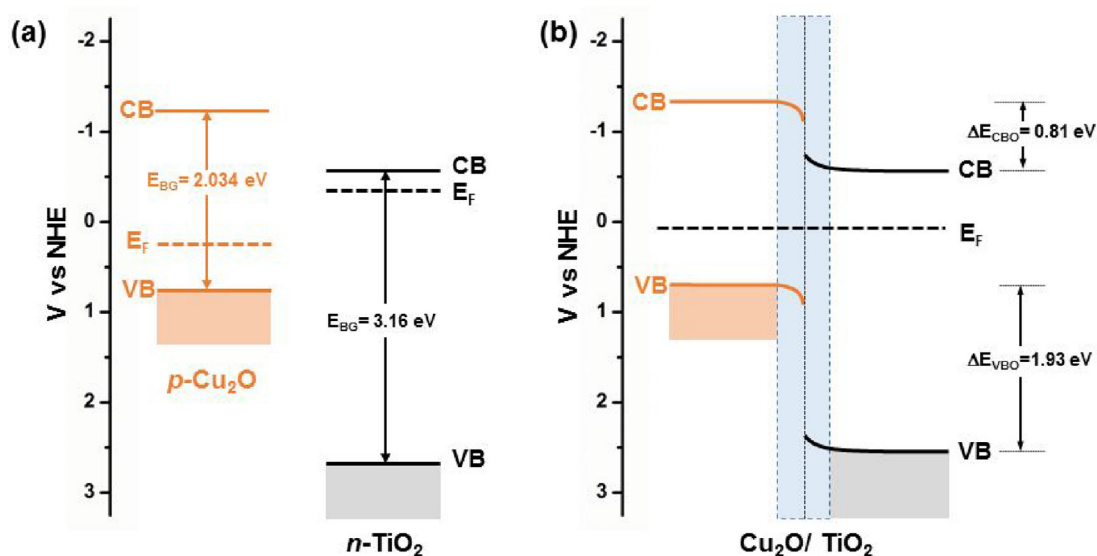
Fig. 3a shows the XPS survey spectrum for  $\text{Cu}_2\text{O}/\text{TiO}_2$  composites, which demonstrates the existence of Cu, Ti and O in the sample. The C 1s peak, associated to the widespread presence of carbon in the environment is also clearly observed. Fig. 3b compares the high resolution XPS spectrum in the region of  $\text{Cu} 2p_{3/2}$  for pure  $\text{Cu}_2\text{O}$  and  $\text{Cu}_2\text{O}/\text{TiO}_2$  composite. The same deconvolution procedure was applied throughout the analysis of the peaks and basically involves the subtraction of a Shirley-type baseline and the use of Voigt-type functions to reproduce the spectra. For  $\text{Cu}_2\text{O}$  the main peak

is centered at 932.01 eV and is readily assigned to Cu(I) [35], while the shake-up satellite peaks with higher binding energy (933.98, 942.46, 943.98 eV) confirm the presence of an unfilled Cu 3d shell corresponding to Cu(II) species at the  $\text{Cu}_2\text{O}$  surface. This observation has been commonly attributed to the oxidation of Cu(I) during sample preparation for analysis [36]. For the  $\text{Cu}_2\text{O}/\text{TiO}_2$  composites, the position of the peaks is similar (See Fig. 3b) but the relative area Cu(I):Cu(II) obtained through the deconvolution of  $\text{Cu} 2p_{3/2}$  spectrum changes from 0.75:0.25 (for pure  $\text{Cu}_2\text{O}$ ) to 0.65:0.35 (in  $\text{Cu}_2\text{O}/\text{TiO}_2$  composites). The previous information together with the results from XRD diffractograms suggests that a minimum fraction of surface copper changes its electronic state during the solothermal treatment.

The high resolution XPS spectrum of Ti 2p in  $\text{Cu}_2\text{O}/\text{TiO}_2$  can be adjusted with great accuracy with two Voigt-type functions centered in 458.31 and 464.19 eV that can be assigned to  $\text{Ti} 2p_{3/2}$  and  $\text{Ti} 2p_{1/2}$ , respectively for Ti(IV) [37] in agreement with the results



**Fig. 4.** (a) UV-visible absorption spectra for solid samples. (b) UPS spectra for Cu<sub>2</sub>O octahedral (top), pure TiO<sub>2</sub> (center), and Cu<sub>2</sub>O/TiO<sub>2</sub> heterostructure (bottom). Extrapolations to the left and right hand sides correspond to the onset values for the valence band and secondary electron spectra, respectively. A complete description is available in the Supporting Information (see Fig. S5).



**Fig. 5.** Energy band diagrams for (a) Cu<sub>2</sub>O and TiO<sub>2</sub> before contact and (b) Cu<sub>2</sub>O/TiO<sub>2</sub> composite.

obtained for pure TiO<sub>2</sub> (Fig. S2). The deconvolution peaks of the O 1s spectrum for the Cu<sub>2</sub>O/TiO<sub>2</sub> requires a greater number of components than for Cu<sub>2</sub>O and in particular, three curves centered in 530.01, 532.08 and 533.63 eV are necessary. The two last components could be assigned to hydroxyl groups and water molecules adsorbed on the TiO<sub>2</sub> surface [38,39]. It should be noticed that the hydroxylation and the presence of adsorbed water molecules benefits CO<sub>2</sub> photoreduction [20] (See Fig. S3, for details).

#### 3.4. Optical properties and energy bands alignments at the heterojunction

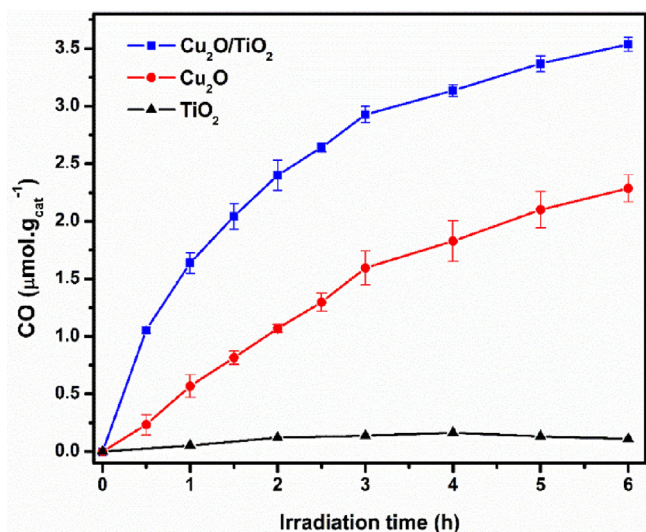
Assessing the bands edge positions of Cu<sub>2</sub>O and TiO<sub>2</sub> and the band alignment is important to understand the electron transfer process at the heterojunction and to determine the reactions that are thermodynamically feasible. The combined information from DRUV, UPS and XPS spectroscopies was used to determine electronic band alignments and the construction of the band energy diagram as discussed below.

DRUV spectra of Cu<sub>2</sub>O, pure TiO<sub>2</sub>, and Cu<sub>2</sub>O/TiO<sub>2</sub> are shown in Fig. 4a. The absorption edges for pure Cu<sub>2</sub>O and TiO<sub>2</sub> samples were approximately 600 and 390 nm, respectively. Analysis of Tauc plots

indicates that the optical bandgaps are E<sub>BG</sub> (Cu<sub>2</sub>O) = 2.03 eV, (as a direct semiconductor) and E<sub>BG</sub> (TiO<sub>2</sub>) = 3.16 eV, (as an indirect semiconductor) [40] (Fig. S4). These figures agree with the previously reported values for Cu<sub>2</sub>O [41] and TiO<sub>2</sub> [10].

Fig. 4b shows the UPS spectra of Cu<sub>2</sub>O, TiO<sub>2</sub> and Cu<sub>2</sub>O/TiO<sub>2</sub>. By applying the method of linear approximation to the UPS spectra, the work function and the corresponding energy of the Fermi level of Cu<sub>2</sub>O were estimated to be 4.70 and −4.70 eV, respectively. Similarly, the valence band maximum was calculated to be −5.20 eV. Considering the average bandgap energy value (2.03 eV for Cu<sub>2</sub>O) obtained from the Tauc plots (Fig. S4, Supporting Information), the minimum of the conduction band is located at −3.17 eV. The quantities previously determined are referred to the vacuum level. Therefore, according to the relationship between the potential of the normal hydrogen electrode (NHE), E\*, and the energy of the vacuum (E<sub>abs</sub>), E<sub>abs</sub> = −E\* − 4.44 (at 298 K), the conduction and valence bands of Cu<sub>2</sub>O are poised at −1.27 and 0.76 eV, respectively. Similarly, following this procedure for TiO<sub>2</sub>, the estimated Fermi level (E<sub>F</sub>), CB and VB vs NHE are −0.38, −0.57 and 2.59 eV, respectively (Fig. S5, Supporting Information).

After contact, the Fermi level of Cu<sub>2</sub>O and TiO<sub>2</sub> equilibrates due to the formation of a *p-n* heterojunction at the interface of the



**Fig. 6.** Carbon monoxide evolution over Cu<sub>2</sub>O/TiO<sub>2</sub> heterostructure (blue), Cu<sub>2</sub>O (red) and TiO<sub>2</sub> (black) under UV-visible irradiation ( $\lambda \geq 305$  nm). (For interpretation of the references to colour in this figure legend, the reader is referred to the web version of this article.)

Cu<sub>2</sub>O/TiO<sub>2</sub> composite [42]. This process alters the band positions, as observed by XPS. Fig. S5 in the Supporting Information provides the XPS spectra. The band alignment at the interface of Cu<sub>2</sub>O/TiO<sub>2</sub> was determined following the method of Kraut [43]. To accurately determine the valence band offset,  $\Delta E_{VBO}$ , the energy difference between the core level ( $E_{CL}$ ) and the valence band maximum ( $E_{VBM}$ ) in the pure materials, as well as the energy difference between the core levels at the interface of the heterostructure ( $\Delta E_{CL}^{Int}$ ) are needed. Eqs. (2) and (3) are used to calculate  $\Delta E_{VBO}$  and  $\Delta E_{CL}^{Int}$ , respectively:

$$\Delta E_{VBO} = (E_{CL}^{Cu_2O} - E_{VBM}^{Cu_2O}) - (E_{CL}^{TiO_2} - E_{VBM}^{TiO_2}) + \Delta E_{CL}^{Int} \quad (2)$$

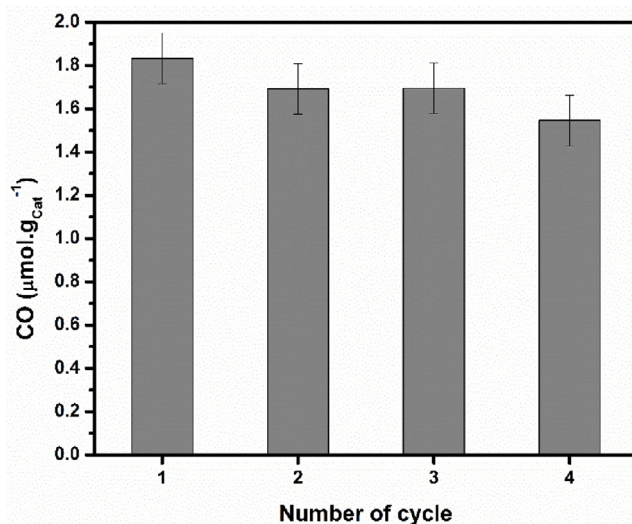
$$\Delta E_{CL}^{Int} = (E_{CL}^{TiO_2} + E_{CL}^{Cu_2O})_{Cu_2O/TiO_2} \quad (3)$$

The conduction band offset,  $\Delta E_{CBO}$ , can be readily obtained from the bandgap energies ( $E_{BG}$ ) of the pure materials and  $\Delta E_{VBO}$ :

$$\Delta E_{CBO} = E_{BG}^{Cu_2O} - E_{BG}^{TiO_2} + \Delta E_{VBO} \quad (4)$$

Combining the information gathered during XPS and DRUV analyses reveals that for the nanocomposite  $\Delta E_{VBO} = 1.93$  eV and  $\Delta E_{CBO} = 0.81$  eV (Fig. 5). The energy difference between the conduction and valence bands for the materials in the composite are about 0.11 eV higher than the values before contact. From the previous observation and the onset value of the composite (Fig. 4b), the valence band maximum of Cu<sub>2</sub>O in the composite is calculated to lie at 0.62 eV (vs NHE). The last figure is useful to calculate the apparent bandgap of Cu<sub>2</sub>O in the composite, which together with  $\Delta E_{VBO}$  and  $\Delta E_{CBO}$  is needed to estimate the minimum of the conduction band of Cu<sub>2</sub>O in the composite to be  $-1.39$  eV (vs NHE) and the valence band maximum and conduction band minimum of TiO<sub>2</sub>, which are depicted at 2.55 and  $-0.58$  eV in Fig. 5, respectively [20].

Fig. 5 shows the energy diagram indicating the formation of a Type II (staggered) band heterostructure, whose highly energetic conduction band favors the reduction of CO<sub>2</sub> through photogenerated electrons. We assume that the assignment of the composite bands corresponds to values that are far from the Cu<sub>2</sub>O/TiO<sub>2</sub> interface. Taking into account this observation, and considering that the energy differences between CB and  $E_F$  for TiO<sub>2</sub> as well as from VB and  $E_F$  for Cu<sub>2</sub>O do not remain constant before and after contact, we showed in Fig. 5 the associated band bendings [44–46].



**Fig. 7.** Carbon monoxide production during 1 h-period in the recycling experiments (see text for details).

The energy difference between the valence band maximum of Cu<sub>2</sub>O and the conduction band minimum of TiO<sub>2</sub> can be estimated as 1.33 and 1.2 eV for the separated semiconductors and the composite, respectively. This change indicates that under UV-visible irradiation, the electron transfer from the conduction band of TiO<sub>2</sub> to the valence band of Cu<sub>2</sub>O (direct Z-scheme) could be favored due to the increased overlap of energy levels involved upon *p-n* heterojunction formation.

### 3.5. Photocatalytic CO<sub>2</sub> reduction

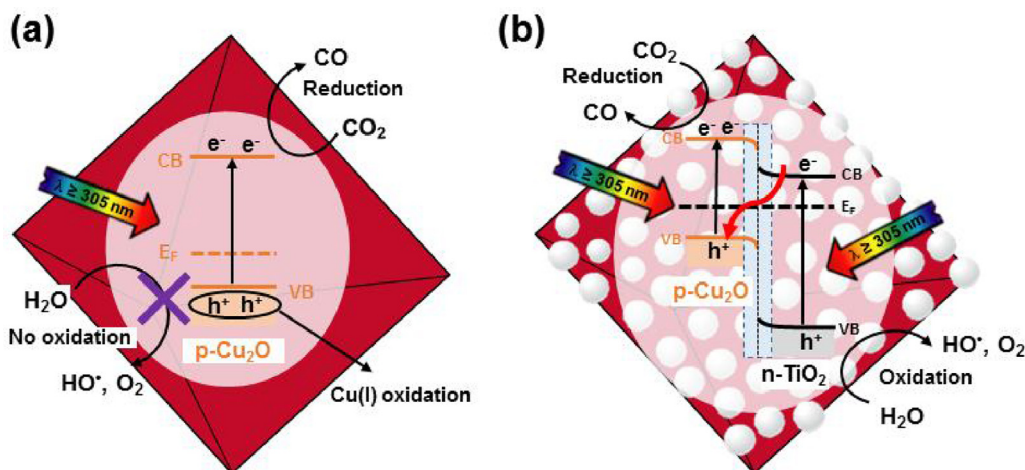
Irradiation of the composite with  $\lambda \geq 305$  nm results in the selective formation of CO, (and O<sub>2</sub>). The time series for CO evolution from the irradiated composite is shown in Fig. 6 together with the results obtained under similar conditions for pure Cu<sub>2</sub>O and TiO<sub>2</sub>. Control experiments for Cu<sub>2</sub>O/TiO<sub>2</sub> showed that no CO evolved in the dark or when the irradiation was carried out in the absence of CO<sub>2</sub> under 1 atm argon (Fig. S6).

It is apparent from the results in Fig. 6 that the formation of the *p-n* heterojunction improves the CO<sub>2</sub> reduction efficiency. Considering the same initial stage of irradiation, the CO evolution rates were calculated to be 2.11 and 0.55  $\mu\text{mol.g}_{\text{cat}}^{-1} \text{h}^{-1}$  for Cu<sub>2</sub>O/TiO<sub>2</sub> and Cu<sub>2</sub>O, respectively, which represents a *ca.* 4-times enhancement for the composite. The ratio from the areas for the convolution of the irradiance of the lamp with the DRUV spectra of 1) Cu<sub>2</sub>O/TiO<sub>2</sub> and 2) Cu<sub>2</sub>O in the interval  $305 \leq \lambda \leq 665$  nm indicates the nanocomposite absorbs *ca.* 2-times more photons than Cu<sub>2</sub>O per unit mass under the experimental conditions in Fig. 6. Thus, only a 50% of the enhanced photocatalytic activity can be explained to arise from an increment in the oscillator strength (or the resulting increase of the integrated absorption per unit mass). Therefore, the synergistic effect from the materials in the nanocomposite is proposed to provide the remaining 50% photocatalytic enhancement.

The ratio of quantum efficiencies for CO production,  $\Phi_{\text{CO}}$ , for Cu<sub>2</sub>O/TiO<sub>2</sub> and Cu<sub>2</sub>O can be easily calculated from the established relationships between the initial reaction rates ( $R_{0,\text{CO}}$ ) and the absorbed photon fluxes  $I_a$ , which was established above,

$$\frac{\Phi_{\text{CO}}^{Cu_2O/TiO_2}}{\Phi_{\text{CO}}^{Cu_2O}} = \frac{R_{0,\text{CO}}^{Cu_2O/TiO_2} / I_a^{Cu_2O/TiO_2}}{R_{0,\text{CO}}^{Cu_2O} / I_a^{Cu_2O}} = \frac{R_{0,\text{CO}}^{Cu_2O/TiO_2}}{R_{0,\text{CO}}^{Cu_2O}} \times \frac{I_a^{Cu_2O}}{I_a^{Cu_2O/TiO_2}} = \frac{4}{1} \times \frac{1}{2} = 2 \quad (5)$$

Therefore, the relative quantum efficiencies in Eq. (5) clearly demonstrate that Cu<sub>2</sub>O/TiO<sub>2</sub> has an enhanced photocatalytic activity over Cu<sub>2</sub>O.



**Scheme 2.** Sketch of the proposed mechanism to account for CO<sub>2</sub> reduction induced by UV–vis irradiation ( $\lambda \geq 305$  nm)) of (a) octahedral Cu<sub>2</sub>O and (b) Cu<sub>2</sub>O/TiO<sub>2</sub> composite.

**Table 1**

Summary of XPS analysis for Cu<sub>2</sub>O and Cu<sub>2</sub>O/TiO<sub>2</sub> photocatalysts at different irradiation times.

Irradiation time (h)	Relative Components		Relative Components	
	Cu <sub>2</sub> O		Cu <sub>2</sub> O/TiO <sub>2</sub>	
	Cu(I)	Cu(II)	Cu(I)	Cu(II)
0	0.75	0.25	0.65	0.35
3	0.69	0.31	0.64	0.36
6	0.60	0.40	0.65	0.35

Another point of interest is that the evolution of CO decreases with the irradiation time. This is a characteristic of many photocatalytic processes [10] and, in particular, has been observed for the photocatalytic CO<sub>2</sub> reduction over  $\alpha$ -Fe<sub>2</sub>O<sub>3</sub>/Cu<sub>2</sub>O [20], and over pure Cu<sub>2</sub>O and Cu<sub>2</sub>O/RuOx systems [47]. The actual reason for this behavior is not well understood but possible causes for the non-linear time profile are commonly ascribed to the strong interaction between Cu<sub>2</sub>O and CO, or the formation of O<sub>2</sub>, which competes with carbon dioxide reduction [47]. To get more insight, we investigate the rate of CO evolution in repeated 1 h irradiation cycles. After each run, the reactor was thoroughly evacuated and re-loaded with the same amount of CO<sub>2</sub> and H<sub>2</sub>O before irradiation. Fig. 7 shows that after 4 cycles, CO yields slightly diminish, in sharp contrast with the results obtained under continuous irradiation. These experiments indicate that the inhibition observed in Fig. 6 is not irreversible, and the catalyst may be recycled without significant loss of efficiency, if the products are periodically removed.

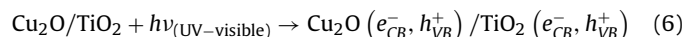
### 3.6. Analysis of the reaction mechanism

As discussed above, two different schemes commonly referred as (A) *double-charge transfer* and (B) *direct Z-scheme mechanisms* are often proposed to explain the electron transfer processes across the interface in *p-n* heterojunctions after the simultaneous excitation of the centers of Type II heterostructures [19,48]. To determine which of these two mechanisms is operative, the photocatalyst was separated from the reactor and analyzed by XPS after different photons doses to monitor the possible change in the copper oxidation state during the irradiation experiments. The most relevant results are summarized in Table 1, which details are provided in Fig. S7.

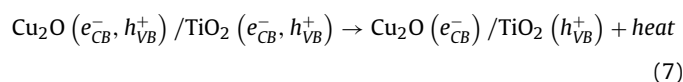
XPS and XRD analysis allow us to discard the possible photoreduction of Cu(I) [30,46] during the irradiation experiments for Cu<sub>2</sub>O and Cu<sub>2</sub>O/TiO<sub>2</sub> composite as no Cu(0) could be observed. On the other hand, as shown in Table 1, XPS analysis of pure Cu<sub>2</sub>O shows

a relative increment of the contribution of Cu(II) in the Cu 2p<sub>3/2</sub> in relation to the Cu(I) content. Cu<sub>2</sub>O photocorrosion might have been anticipated since Cu<sub>2</sub>O irradiation generates a highly reductive conduction band electron,  $E_{CB} = -1.27$  eV, but a valence band hole with insufficient energy to oxidize water [11] (Scheme 2a). This key point is also demonstrated during experiments presented below that have quantified the production of HO• radicals.

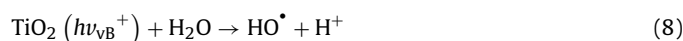
By contrast, in the Cu<sub>2</sub>O/TiO<sub>2</sub> composite, the ratio Cu(II)/Cu(I) keeps constant along the entire photocatalytic reaction. To account for these results, we proposed that UV-visible irradiation induces the formation of an electron/hole pair in each photocatalytic center, Eq. (6):



and after carrier generation, the electrons in the conduction band of TiO<sub>2</sub> are transferred to Cu<sub>2</sub>O (see Scheme 2b):



and the TiO<sub>2</sub> holes localize on surface oxygen, forming  $(\text{Ti}^{3+})_{\text{latt}}$  and  $(\text{Ti}^{4+}\text{O}^{\bullet-})_{\text{surf}}$  paramagnetic species, respectively. By these processes, we achieved a heterostructure with strongly reductive electrons,  $\text{Cu}_2\text{O}(e_{CB}^-)$ , and oxidative holes,  $\text{TiO}_2(h_{VB}^+)$ , able to oxidize water. Water oxidation by TiO<sub>2</sub> is not only thermodynamically allowed but also kinetically feasible as revealed by EPR experiments, which indicate that surface trapped holes react with adsorbed molecules even at the extremely low temperatures used in the EPR studies [49].

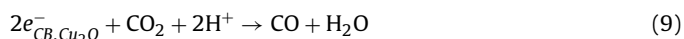


The fluorescence spectra for the production of 7-hydroxycoumarin during irradiation of Cu<sub>2</sub>O/TiO<sub>2</sub> is presented in Fig. S8 (Supplementary Information). The characteristic fluorescence peak at  $\lambda_{\text{em}} = 456$  nm matches the reported spectrum of 7-hydroxycoumarin [28], which implies that only this species is produced [27]. Importantly, no peak for 7-hydroxycoumarin production was observed during 1) a dark control with Cu<sub>2</sub>O/TiO<sub>2</sub>, or 2) irradiated Cu<sub>2</sub>O. These controls strongly support the assignment of HO•, an important intermediate for O<sub>2</sub> production, as a unique product from the photoactivity of the nanocomposite. The linear relationship between fluorescence intensity and irradiation time (Fig. S8) indicates a constant production rate for 7-hydroxycoumarin. Consequently, the production of HO• displayed in the inset of Fig. S8 (Supplementary Information) is

quantified by standard addition of 7-hydroxycoumarin to the samples. From the slope of the inset in Fig. S8 (Supplementary Information), the production rate of HO• is  $0.22 (\pm 0.03) \mu\text{mol g}_{\text{cat}}^{-1} \text{h}^{-1}$ . Comparatively, after 3 h of irradiation, the rate of production of CO is 5-times larger. By correcting the previous rates in a per photon basis using the stoichiometric consumptions of  $2 e_{\text{CB}}^-$  and  $1 h_{\text{VB}}^+$  for CO and HO• species produced, respectively, indicates that 40% of the theoretical amount of HO• is trapped by 7-hydroxycoumarin. The previous observation is not surprising for HO•, a high reactivity and short lived species with a lifetime  $\tau \sim 1 \text{ ns}$  [28].

In the presence of CO<sub>2</sub>, additional evidence of water oxidation by Cu<sub>2</sub>O/TiO<sub>2</sub> nanostructures, comes from the analysis of the ratio between CO and O<sub>2</sub> yields. After 1 h of irradiation, we determined  $\text{CO}/\mu\text{mol g}_{\text{cat}}^{-1} = 1.69 \pm 0.12$  and  $\text{O}_2/\mu\text{mol g}_{\text{cat}}^{-1} = 0.83 \pm 0.06$ , which gives  $\text{CO}/\text{O}_2 = 1.69/0.83 = 2.03$ , in close agreement with the expected reaction stoichiometry.

According to the Z-scheme, the anodic reaction takes place in the Cu<sub>2</sub>O surface. In this regard, it is interesting to notice that the one electron homogeneous reduction potential for the monoelectronic reduction of CO<sub>2</sub> (Eq. (1)), is more negative than the conduction band level of Cu<sub>2</sub>O ( $E_{\text{CB}}(\text{Cu}_2\text{O}) = -1.39 \text{ eV}$ ). Thus, in order to make the CO<sub>2</sub> reduction thermodynamically feasible, some stabilization must be invoked [50,51]. Alternatively, rapid hole scavenging by adsorbed molecular water could lead to electron accumulation in Cu<sub>2</sub>O, and reduce CO<sub>2</sub> by a two-electron process,



with lower energy demand ( $E^*(\text{CO}_2/\text{CO}) = -0.53 \text{ V}$ ) [52], avoiding the formation of the high energy intermediate,  $\text{CO}_2^{\bullet-}$ . However, current evidence does not allow distinguishing between the two possibilities.

#### 4. Conclusion

A Cu<sub>2</sub>O/TiO<sub>2</sub> photocatalyst with high morphologic and crystallographic quality was synthesized by a simple solvothermal method. The new material shows a 4-times enhancement for the photoreduction rate of CO<sub>2</sub> induced by UV-visible irradiation in comparison to pure Cu<sub>2</sub>O. Detailed analyses of the photocatalysts by XPS spectroscopy under operation, as well as the formation of HO• radicals in the Cu<sub>2</sub>O/TiO<sub>2</sub> system, at variance with the results for pure Cu<sub>2</sub>O, unequivocally prove a Z-scheme mechanism. The results obtained are of practical interest for photocatalysis and are also relevant for the understanding of photoinduced interfacial charge-transfer processes at the *p-n* heterojunction.

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#### Appendix A. Supplementary data

Supplementary data associated with this article can be found, in the online version, at <http://dx.doi.org/10.1016/j.apcatb.2017.05.058>.

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